

Effects of Thickness on Optical and Structural Properties of Lead Sulphide (PbS) Thin Film Prepared by Chemical Bath Deposition (CBD)

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Abstract

Nanocrystalline PbS thin films were deposited on glass substrates with various deposition times using chemical bath deposition technique. PbS thin films were prepared by using a solution of Lead nitrate and thiourea by using NaOH salt as a complexing agent and ammonia solution as a pH adjuster at 65°C. The study was carried out for thickness in the range of (0.16-1.03µm). The structural properties were carried out by X-ray diffraction (XRD). The X-ray diffraction patterns reveal that the films exhibit the cubic face centered structure. The crystalline sizes of the films increased with increasing film thickness. The morphological properties were studied by SEM and showed that films were well adherent to the substrate, uniform, small crystal size and covered the entire substrate surface completely. The optical properties of these films have been studied and show that PbS thin films have the values of energy gap varied between (0.88-1.34 eV) with increasing film thickness.

Keywords: CBD, XRD, band gap, PbS thin film

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Introduction

Lead Sulphide (PbS) is cluster IV-VI chalcogenides semiconductor that has direct optical band gap of 0.41 electron volt and has a cubic structure [1,21]. The colour of PbS is grayish-black. It's high absorbance within the UV-Vis spectral vary and high transmittance within the infrared spectral region of allowing its use in infrared detectors [2,9]. Thin film of lead sulphide is useful in the manufacture of transistors[3], photoconductive infra-red detectors[4], windows, contact rectifiers, prisms, lenses[5], photovoltaic application in solar cells[6], solar control coating[7], optoelectronic devices[8], antireflection coatings and for solar thermal applications[9] in flat-plate collectors, house heating for solar chick brooding (11), Photo resistance, diode lasers, humidity and temperature sensors, decorative coatings and solar control coatings [12], pollution monitor[25].

Many techniques have been reported for the deposition of thin films such as Chemical spray deposition (CSP) [26], Molecular beam epitaxy [13], chemical vapor deposition (CVD) [15], sol-gel spin coating [16], successive ion layer absorption and reaction (SILAR) method[4], chemical bath deposition (CBD) [13] ect. Among them chemical bath deposition (CBD) method does not require sophisticated instrumentation [17]. It is relatively cheap, convenient for large area deposition [11], simple to handle, and capable of yielding good quality thin films [17].

The film thickness being one of parameters, which affects the properties of the thin film. The film thickness can be measured by the gravimetric method [18], Tolansky method [12], interference fringes, ellipsometry method and weight difference method [18].

In this research Lead sulphide (PbS) thin film were deposited at 65°C from lead nitrate ($Pb(NO_3)_2$) thiourea ($SC(NH_2)_2$), sodium hydroxide (NaOH) and ammonium hydroxide (NH_4OH) by using Chemical bath deposition method and it's thickness were measured by weight difference method

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(Pb(NO₃)₂) thiourea (SC(NH₂)₂), caustic soda (NaOH) and ammonia (NH₄OH) by using Chemical bath deposition method and it's thickness were measured by weight difference method

Materials and Methods

Lead sulphide thin films were deposited on substrate by the chemical bath deposition method. The precursor chemicals were lead nitrate (Pb(NO₃)₂), thiourea (SC(NH₂)₂), sodium hydroxide (NaOH), ammonium hydroxide (NH₄OH). Thiourea is used as our sulphide ion source and leads nitrate as our lead ion source. Ammonia is as PH adjuster, NaOH as complex agent. Lead sulphide were constituted from a solution of 5 ml of 0.06 M lead nitrate (Pb(NO₃)₂), 5 ml of 0.2 M thiourea, 10 ml of 0.25M sodium hydroxide, 7ml ammonia and distilled water of 43ml were added to the solution making it total of 70 ml. Cleaned substrates were vertically immersed into the solution and was maintained 65°C temperature. After the deposition, the slides were removed from the chemical bath at 17,22,27,32min for four samples respectively. the thin film was rinsed with distilled water and dried in air. The resulting films were homogeneous, uniform and well adhered to the substrate with dark surface. A mirror-like surface was obtained. The thickness was measured by weight different method. In this case the mass of cleaned substrate was measured before deposition and after deposition. Then the difference between two mass was obtained by subtracting the mass of substrate before deposition from mass of substrate with thin film. From its general properties, the density of PbS is constant which is 7.60 g/cm³. The length and width of deposited thin film were measured by ruler. Then the area of the film was calculated from width and length of the film. The Thickness was calculated by using the following formula [18].

$$T = \frac{\Delta M}{\rho M} \quad (1)$$

where A is area of the film, ρ is density of the film, difference between two mass and T is thickness of the film. Four samples were measured according to the following table.

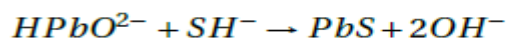
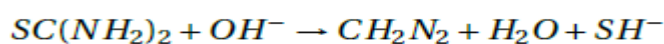
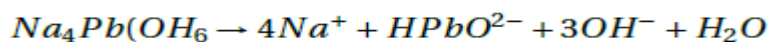
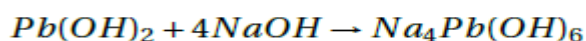
Table 1: Thickness of measured each film

time(min)	A(cm) ²	m. B deposition(μg)	m.A deposition(μg)	M.D(μg)	thickness(μM)
17	5	42708	42714	6	0.16
22	5	42162	42181	19	0.5
27	5	45623	45660	37	0.97
32	5	44929	44968	39	1.03

where t,m.B,m.A,m.D are time,mass before deposition,mass After deposition,mass difference respectively

Result and discussion

Reaction mechanism



[21]

Optical characterization

The optical energy gap (E_g) of PbS thin films were determined using Uv-3600. The optical absorption spectrum of PbS were calculated in wave length range of 700-1800. The optical band gap was calculated by using stern equation [20,24].

$$A = \frac{[K(h\nu - E_g)]^{\frac{n}{2}}}{h\nu} \quad (2)$$

Where ν is the frequency of the incident photon, h is Planck's constant A is constants and n is the number which characterizes the optical processes. The plots of the (Ahν)² as a function of hν are shown in fig 1. in accordance with equation 2. The film thickness dependence of band gap energy can be due to factors like the existence of amorphous phases in thin films, Quantum confinement effect [6,19] and change in barrier height owing to change in grain size in polycrystalline films [6].

In this case, it can be considered that the thickness dependence of band gap energy is concerned with the crystallinity of the films. The decrease in band gap energy with film thickness is similar to its crystallite size dependence, which indicates the decrease in barrier height at the boundaries with an increase in crystallite size. The value of energy was decrease as thickness of the film is increase. This may be due to increase in crystallite size, crystal imperfection and, decrease in stacking faults [22]. The larger crystallite size leads to an increase in absorption. As a result, the optical absorption edge shift towards longer wavelengths and consequently band gap decreases. As we see from fig.1. the value of band gap is 1.34, 1.12, 0.92, 0.88 eV. for thickness 0.16,0.5,0.97,1.03 μm respectively.

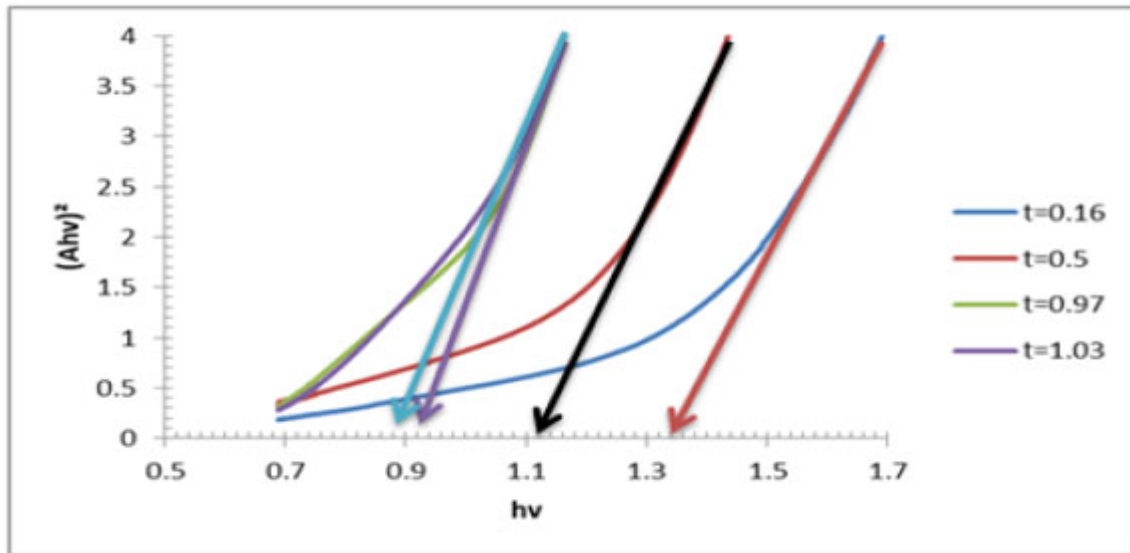


Fig1. plot of $(Ahv)^2$ versus band gap for all samples

Structural Characterization

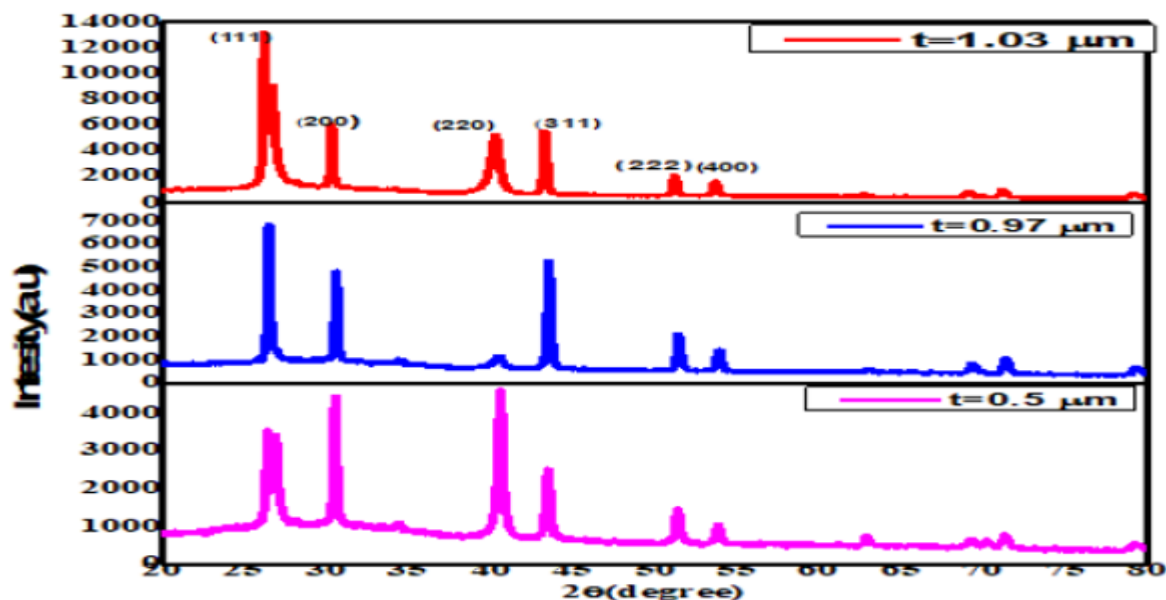


Figure 2. XRD characterization for all 0.5,0.97 and 1.03 μm

From Fig.2. the X-ray diffraction of PbS thin films with thickness 0.5, 0.97 and 1.03 μm . For thickness 0.5 μm the diffraction peaks at 26.450,30.540, 40.600, 43.480, 51.420 and 53.850 which, respectively, correspond to the (111), (200), (220),(311), (222) and (331), Miller planes of the PbS phase with cubic structure (galena), for thickness 0.97 μm produced peaks at 26.50°, 30.60°, 43.56°, 51.46°,53.91° and 71.39° corresponding to the Miller planes (111), (200), (220), (311),(222) and (400) respectively as well as the thickness 1.03 μm produced peaks at 26.23°, 30.34°, 40.30°, 43.32° and 51.24° corresponding to the Miller planes (111), (200), (220), (311),(222) and (400). The diffraction pattern was well matched with the standard JCPDS data file reference code: No.00-005-

0592 of the face centered cubic structure. The intensity of the peaks in 1.03 μm films is higher than those of the 0.5 and 0.97 μm films because the thickness of the films of 1.03 μm was larger than others films. The preferred growth orientation in films changed from (220) to (111) with increasing thickness. This due to formation of Nano crystalline film. The presence of well-defined and sharp peaks in all the XRD patterns is attributed to a high degree of crystallinity in the thin films. The crystallite size can be estimated by Debye-Scherrer's equation using the line broadening of the highest intensity peak (220).

$$D = \frac{0.94\lambda}{\beta \cos \theta} \quad (3)$$

Table 2: plot of crystalline size versus thickness of PbS Thin film

$2\theta(\text{deg})$	Thickness(μm)	Time(min)	hkl	FWHM(rad)	D(nm)	$(\epsilon) \text{ lines}^{-2}\text{m}^{-4}$	$(\delta)(\hat{t}) \text{ lines}/\text{m}^2$
40.6019	0.5	22	220	0.41396	0.06	3.65×10^{-3}	0.278×10^{15}
26.5089	0.97	27	111	0.19787	0.126	2.75×10^{-3}	0.063×10^{15}
26.23251	1.03	32	111	0.18692	0.133	2.60×10^{-3}	0.057×10^{15}

Where, λ is the X-ray wavelength (1.5406\AA), (β) is the line broadening of diffraction at half of its maximum intensity, k is the constant numerator value (0.94). the lattice parameters for cubic phase(hkl) calculated as follows. The strain and dislocation density (δ) were evaluated by using Williamson and tangent relations, respectively

$$E = \frac{\beta \cos \theta}{4} \quad (4)$$

$$\delta = \frac{1}{D^2} \quad (5)$$

The obtained dislocation density and macrostrain values were low, which indicates the presence of minimum defects in the prepared films. The strain decreased with increasing film thickness indicating that the improvement of the stoichiometric ratio of PbS thin films with increasing film thickness. The dislocation density also decreased with increasing crystallite size. It may be due to a decrease in the occurrence of grain boundaries because of an increase in the crystallite size of the film with increasing film [23]. FWHM of the peaks decreased with increase in thicknesses of PbS thin films, thus confirming inverse relation between dislocation density and crystallinity and hence it is concluded that increase in thickness improved crystallinity. As the thickness increased, the crystallite size changed, so dislocation density and strain decreases. The lattice constant, a_{hkl} for PbS thin films with the cubic face centered structure was calculated using Bragg diffraction equations:

$$a_{hkl} = d_{hkl} \sqrt{h^2 + k^2 + l^2} \quad (6)$$

Where d_{hkl} represents the distance between the (hkl) atomic lattice planes. The lattice constant of the PbS thin films varies as 6.2\AA , 4.5\AA and 2.5\AA with increasing thickness. where δ is dislocation density, ϵ is strain size.

Morphological Characterization

The morphology of PbS thin films of thickness 0.97 μm deposited on substrate was examined by scanning electron microscopy (SEM) (Fig.3). The surface of the film is uniform and covers the entire substrate without cracks and holes. And also, it reveals that the film formed is smooth and dense. the grains in the film are relatively good, compact, homogeneous and uniformly covered at thickness 0.97 μm .

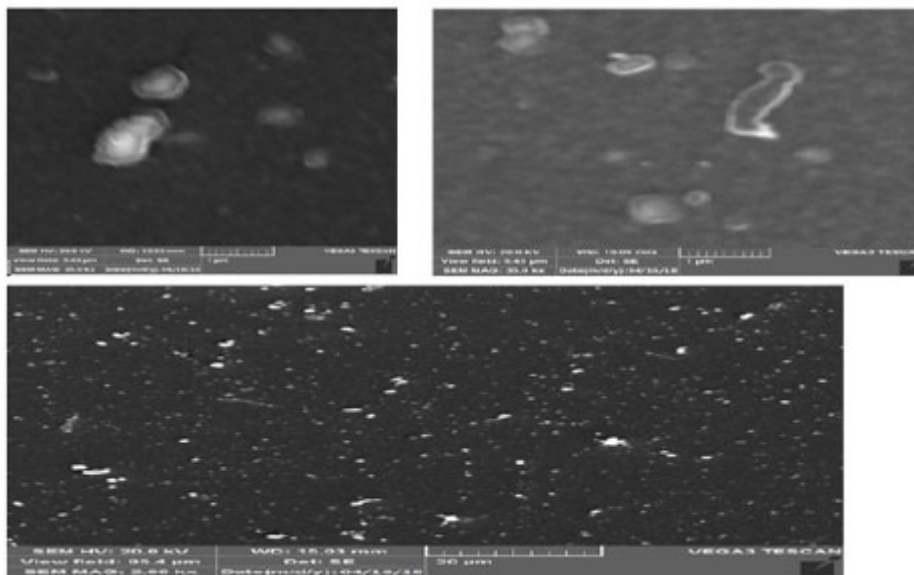


Figure 3: SEM Characterization for thickness 0.97 μ m

Elemental Characterization

The a) and b) show distribution Pb and S respectively. Figure c) shows the EDX analysis of PbS thin Film indicate the presence lead and Sulphur. The calculated Pb/S atomic ratio of thin film deposited at thickness 0.97 μ m was 0.53:0.47 which show the lead was reach than sulphide. The average atomic percentage of S to Pb was 1.13. which showed that the sample was non stoichiometric. but the strong peak is from silicon. This is may be due to the formation of very thin film. Other elements such as Carbon, Oxygen, Ca, and Si are observed. Those are may due to the composition of the silica glass substrate. (The oxygen content found in the EDX spectrum could either emanate from the substrate or due to the films getting oxidized on exposure to the atmosphere).

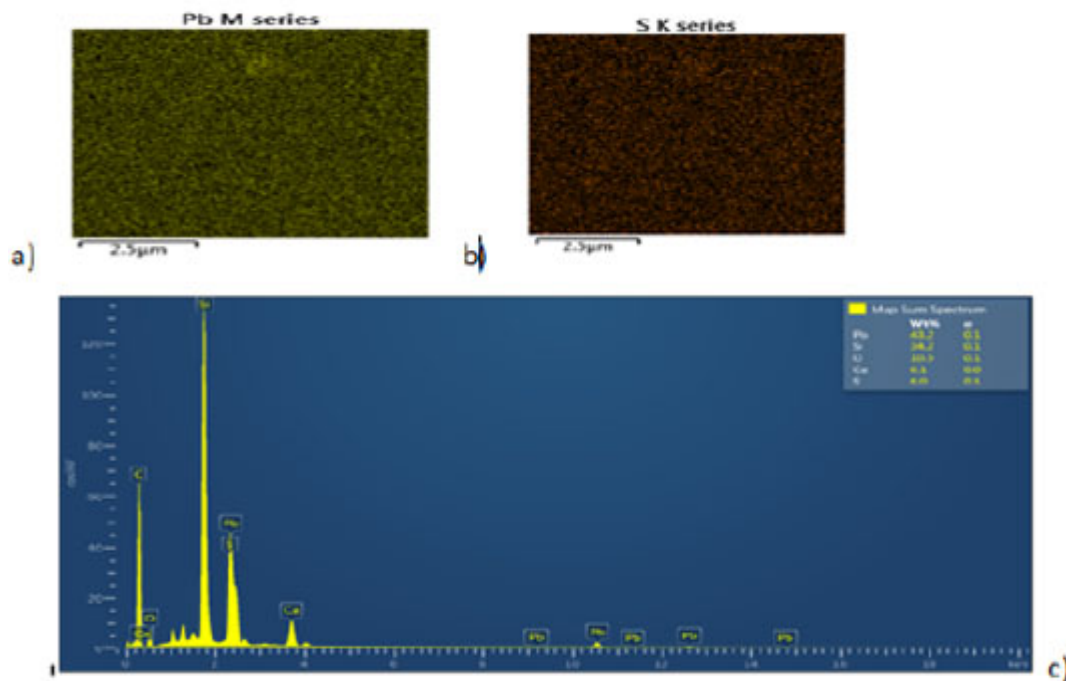


Figure 4: EDS analysis for thickness 0.97 μ m

Conclusion

In this studies PbS thin films at 65⁰C temperature were prepared by a simple chemical Bath Deposition method. The thickness of the films varied from 0.16 μ m to 1.03 μ m. The Effects of thickness on the structural and optical

properties of the films were studied. The characteristics of PbS thin films were investigated by various tools. The X-ray diffraction patterns confirm the proper phase formation of the PbS thin film. For the first sample, the preferential orientation value of (220) plane has the highest value compared to other planes and for second and third samples the preferential orientation to change to (111). The crystalline size of the PbS films increased with increasing film thickness. The values of crystallite size were found to be in the range of 0.6-1.3 nm. The SEM study show that the film is well adherent to the substrate, uniform and covers entire substrate. The results of XRD and SEM show that the deposited PbS film consists of nano-sized crystalline and the crystalline size increases with increasing film thickness. The optical band gap energy varies from 0.86 to 1.34 eV with increasing thickness.

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